# Chemistry for Life, Chemistry for better Life



# **Theoretical Test**



# 2006. 7. 7 Gyeongsan, Korea

#### **General Directions**

- Write your name and code number on each page of the answer sheet.
- You have 5 hours to finish the task. Failure to stop after the STOP command may result in zero points for the task.
- Write answers and calculations within the designated box.
- Use only the pen and the calculator provided.
- There are 30 pages of Problems and 5 pages of Answer Sheet.
- An English-language version is available.
- You may go to the restroom with permission.
- After finishing the examination, place all sheets including Problems and Answer Sheet in the envelope and seal.
- Remain seated until instructed to leave the room.

1 H 1.01	Periodic Table of Elements       2         with atomic masses       4.00									2 He 4.00							
3 Li 6.94	4 Be 9.01											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 19.00	10 Ne 20.18
11 Na 22.99	12 Mg 24.31											13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.07	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 <u>Mn</u> 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.39	31 Ga 69.72	32 Ge 72.61	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc 98.91	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.90	54 Xe 131.29
55 Cs 132.91	56 Ba 137.3	57-71	7 <b>2</b> Hf 178.49	7 <b>3</b> Ta 180.95	7 <b>4</b> W 183.84	75 Re 186.21	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.97	80 Hg 200.59	81 Tl 204.38	82 Pb 207.19	83 Bi 208.98	84 Po 208.98	85 At 209.99	86 <u>Rn</u> 222.02
87 Fr 223	88 Ra 226	89-103	104 Rf 261	105 Db 262	106 Sg 263	107 Bh 264	108 Hs 265	109 Mt 268		•	•	•	•	•	•	•	

]

57	58	59	60	61	62	63	64	65	66	67	68	69	70	71
La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Но	Er	Tm	Yb	Lu
138.91	140.12	140.91	144.24	144.92	150.36	151.96	157.25	158.93	162.50	164.93	167.26	168.93	173.04	174.97
89	90	91	92	93	94	95	96	<b>9</b> 7	98	99	100	101	102	103
Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr
227	232	231	238	237	244	243	247	247	251	252	257	258	259	262

## Constants and useful formulas

Gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ Faraday constant $F = 96485 \text{ C mol}^{-1}$ Use as standard pressure: $p = 1.013 \cdot 10^5 \text{ Pa}$ Use as standard temperature: $T = 25^{\circ}\text{C} = 298.15 \text{ K}$ Avogadro's number $N_A = 6.022 \cdot 10^{23} \text{ mol}^{-1}$ Planck constant $h = 6.626 \cdot 10^{-34} \text{ J s}$ Speed of light $c = 3.00 \cdot 10^8 \text{ m s}^{-1}$ 

$$\Delta G = \Delta H - T\Delta S \qquad \Delta G = -nFE$$
  

$$\Delta G^{0} = -RT \cdot lnK \qquad \Delta G = \Delta G^{0} + RT \cdot lnQ \qquad \text{with } Q = \frac{product \ of \ c(products)}{product \ of \ c(reactands)}$$
  

$$\Delta H(T_{I}) = \Delta H^{0} + (T_{I} - 298.15 \text{ K}) \cdot C_{p} \qquad (C_{p} = \text{constant})$$

Arrhenius equation  $k = A \cdot e^{-\frac{E_o}{R \cdot T}}$ Ideal gas law pV = nRTNernst equation  $E = E^0 + \frac{RT}{nF} \cdot ln \frac{c_{ox}}{c_{red}}$ Beer- Lambert Law  $A = \log \frac{P_0}{P} = \varepsilon \cdot c \cdot d$ V(cylinder) =  $\pi r^2 h$ 

A(sphere) =  $4\pi r^2$ V(sphere) =  $\frac{4}{3}\pi r^3$ 

1 J = 1 N m	1 N = 1 kg m s <sup>-2</sup>	1 Pa = 1 N m <sup>-2</sup>
$1 \text{ W} = 1 \text{ A V} = 1 \text{ J s}^{-1}$	1 C = 1 A s	

#### 1. Avogadro's number

Spherical water droplets are dispersed in argon gas. At  $27^{\circ}$ C, each droplet is 1.0 micrometer in diameter and undergoes collisions with argon. Assume that inter-droplet collisions do not occur. The root-mean-square speed of these droplets was determined to be 0.50 cm/s at  $27^{\circ}$ C. The density of a water droplet is 1.0 g/cm<sup>3</sup>.

1-1. Calculate the average kinetic energy (mv<sup>2</sup>/2) of this droplet at 27°C. The volume of a sphere is given by (4/3)  $\pi$  r<sup>3</sup> where r is the radius.

If the temperature is changed, then droplet size and speed of the droplet will also change. The average kinetic energy of a droplet between  $0^{\circ}$ C and  $100^{\circ}$ C as a function of temperature is found to be linear. Assume that it remains linear below  $0^{\circ}$ C.



At thermal equibrium, the average kinetic energy is the same irrespective of particle masses (equipartition theorem).

The specific heat capacity, at constant volume, of argon (atomic weight, 40) gas is 0.31 J  $g^{-1}$  K<sup>-1</sup>.

1-2. <u>Calculate</u> Avogadro's number without using the ideal gas law, the gas constant, Boltzmann's constant).

## 2. Detection of Hydrogen

Hydrogen is prevalent in the universe. Life in the universe is ultimately based on hydrogen.

2-1. There are about 10<sup>23</sup> stars in the universe. Assume that they are like our sun (radius, 700,000 km; density, 1.4 g/cm<sup>3</sup>; 3/4 hydrogen and 1/4 helium by mass). <u>Estimate</u> the number of stellar protons in the universe to one significant figure.

In the 1920s, Cecilia Payne discovered, by spectral analysis of starlight, that hydrogen is the most abundant element in most stars.

- 2-2. The electronic transition of a hydrogen atom is governed by  $\Delta E(n_i \rightarrow n_f) = -C(1/n_f^2 1/n_i^2)$ , where *m* and *n* are principle quantum numbers, and C is a constant. For detection of the  $\Delta E(3\rightarrow 2)$  transition (656.3 nm in the Balmer series), the electron in the ground state of the hydrogen atom needs to be excited first to the n = 2 state. <u>Calculate</u> the wavelength (in nm) of the absorption line in the starlight corresponding to the  $\Delta E(1\rightarrow 2)$  transition.
- 2-3. According to Wien's law, the wavelength ( $\lambda$ ) corresponding to the maximum light intensity emitted from a blackbody at temperature *T* is given by  $\lambda T = 2.9 \times 10^{-3}$  m K. <u>Calculate</u> the surface temperature of a star whose blackbody radiation has a peak intensity corresponding to the *n* = 1  $\rightarrow$  *n* = 2 excitation of hydrogen.

The ground state of hydrogen is split into two hyperfine levels due to the interaction between the magnetic moment of the proton and that of the electron. In 1951, Purcell discovered a spectral line at 1420 MHz due to the hyperfine transition of hydrogen in interstellar space.

2-4. Hydrogen in interstellar space cannot be excited electronically by starlight. However, the cosmic background radiation, equivalent to 2.7K, can cause the hyperfine transition. <u>Calculate</u> the temperature of a blackbody whose peak intensity corresponds to the 1420 MHz transition.

2-5. Wien generated hydrogen ions by discharge of hydrogen gas at a very low pressure and determined the e/m value, which turned out to be the highest among different gases tested. In 1919, Rutherford bombarded nitrogen with alpha-particles and observed emission of a positively charged particle which turned out to be the hydrogen ion observed by Wien. Rutherford named this particle the "proton". <u>Fill in</u> the blank.

#### 3. Interstellar Chemistry

Early interstellar chemistry is thought to have been a prelude to life on Earth. Molecules can be formed in space via heterogeneous reactions at the surface of dust particles, often called the interstellar ice grains (IIGs). Imagine the reaction between H and C atoms on the IIG surface that forms CH. The CH product can either desorb from the surface or further react, through surface migration, with adsorbed H atoms to form  $CH_2$ ,  $CH_3$ , etc.

Depending on how energetically a molecule "jumps" from its anchored site, it either leaves the surface permanently (desorption) or returns to a new position at the surface (migration). The rates of desorption and migratory jump follow the Arrhenius formula,  $k = A \exp(-E/RT)$ , where *k* is the rate constant for desorption or migratory jump, *A* the jumping frequency, and *E* the activation energy for the respective event.

- **3-1.** Desorption of CH from the IIG surface follows first-order kinetics. <u>Calculate</u> the average residence time of CH on the surface at 20 K. Assume that  $A = 1 \times 10^{12} \text{ s}^{-1}$  and  $E_{\text{des}} = 12 \text{ kJ mol}^{-1}$ .
- **3-2.** Consider the shortest time it would take for one CH unit to move from its initial position to the opposite side of an IIG by successive migratory jumps. Assume that the activation energy for migration ( $E_{mig}$ ) is 6 kJ mol<sup>-1</sup>, and the IIG is a sphere with a 0.1 µm radius. Each migratory jump laterally advances the molecule by 0.3 nm. Show work and <u>choose</u> your answer from (a)-(e) below.

(a) 
$$t \le 1$$
 day (b) 10 day  $\le t \le 10^2$  yr (c)  $10^3$  yr  $\le t \le 10^6$  yr (d)  $10^7$  yr  $\le t \le 10^{10}$  yr (e)  $t \ge 10^{11}$  yr

**3-3.** Consider the reaction of CO with  $H_2$  to form  $H_2$ CO. The activation energy on a metal catalyst is 20 kJ mol<sup>-1</sup>, which produces formaldehyde

at a rate of 1 molecule/s per site at 300 K. <u>Esitmate</u> the rate of formaldehyde formation per site if the reaction takes place at 20 K.

**3-4.** Which is a set of all true statements? <u>Circle</u> one.

(a) Most CH species desorb from the IIG surface before encountering other reactants by surface migration.

(b) IIGs can assist transformation of simple molecules to more complex ones in interstellar space.

(c) For a reaction on the IIG to occur at an appreciable speed during the age of the Universe (1 x  $10^{10}$  yr), the reaction energy barrier must be absent or negligible.

(a) (b) (c) (a, b) (a, c) (b, c) (a, b, c)

# 4. The Chemistry of DNA

4-1. In 1944 Oswald Avery isolated a genetic material and showed, by elemental analysis, that it was a sodium salt of deoxyribonucleic acid. A segment of DNA with formula mass of 1323.72 is shown.



Assuming that equimolar amounts of the four bases are present in DNA, <u>write</u> the number of H atoms per P atom. <u>Calculate</u>, to 3 significant figures, the theoretical weight percentage of H expected upon elemental analysis of DNA.

4-2. Chargaff extracted the separated bases and determined their concentrations by measuring UV absorbance. The Beer-Lambert law was used to obtain the molar concentration. Chargaff discovered the following molar ratio for bases in DNA:

adenine to guanine = 1.43	thymine to cytosine = 1.43
adenine to thymine = 1.02	guanine to cytosine = 1.02

Chargaff's discovery suggested that the bases might exist as pairs in DNA. Watson and Crick mentioned in their celebrated 1953 paper in *Nature*: "It has not escaped our notice that the specific pairing we have postulated immediately suggests a possible copying mechanism for the genetic material."

<u>Draw</u> structures of the specific pairing found in DNA. <u>Indicate</u> hydrogen bonds. Omit the sugar-phosphate backbone.

- 4-3. Mutation can occur through base pairings different from the above. <u>Draw</u> structures of any three alternative base pairs.
- 4-4. The plausibility of the formation of purine and pyrimidine bases in the prebiotic atmosphere of the Earth from HCN, NH<sub>3</sub>, and H<sub>2</sub>O has been demonstrated in the laboratory. <u>Write</u> the minimum number of HCN and H<sub>2</sub>O molecules required for formation of the following compounds.



#### 5. Acid-Base Chemistry

- **5-1.** <u>Calculate</u> [H<sup>+</sup>], [OH<sup>-</sup>], [HSO<sub>4</sub><sup>-</sup>], and [SO<sub>4</sub><sup>2-</sup>] in a 1.0 x 10<sup>-7</sup> M solution of sulfuric acid ( $K_w = 1.0 \times 10^{-14}$ ,  $K_2 = 1.2 \times 10^{-2}$  at 25°C). In your work you may use mass- and charge-balance equations. Answer with two significant figures.
- **5-2.** <u>Calculate</u> the volume of 0.80 M NaOH solution that should be added to a 250 mL aqueous solution containing 3.48 mL of concentrated phosphoric acid in order to prepare a pH 7.4 buffer. Answer with three significant figures. (H<sub>3</sub>PO<sub>4</sub> (aq), purity = 85 % wt/wt, density = 1.69 g/mL, FW = 98.00) (p $K_1$  = 2.15, p $K_2$  = 7.20, p $K_3$  = 12.44).
- **5-3.** The efficacy of a drug is greatly dependent on its ability to be absorbed into the blood stream. Acid-base chemistry plays an important role in drug absorption.

Membrane  
Stomach  

$$pH = 2.0$$
  
 $H^+ + A^- \longrightarrow HA$   
 $HA \longrightarrow H^+ + A^-$ 

Assume that the ionic form (A<sup>-</sup>) of a weakly acidic drug does not penetrate the membrane, whereas the neutral form (HA) freely crosses the membrane. Also assume that equilibrium is established so that the concentration of HA is the same on both sides. <u>Calculate</u> the ratio of the total concentration ([HA] + [A<sup>-</sup>]) of aspirin (acetylsalicylic acid, pK = 3.52) in the blood to that in the stomach.

## 6. Electrochemistry

Water is a very stable molecule, abundant on earth and essential for life. As such, water was long thought to be a chemical element. However, soon after the invention of a voltaic cell in 1800, Nicholson and Carlyle decomposed water into hydrogen and oxygen by electrolysis.

- **6-1.** Water can be thought of as hydrogen oxidized by oxygen. Thus, hydrogen can be recovered by reduction of water, using an aqueous solution of sodium sulfate, at a platinum electrode connected to the negative terminal of a battery. The solution near the electrode becomes basic. <u>Write</u> a balanced half-reaction for the reduction of water.
- **6-2.** Water can also be thought of as oxygen reduced by hydrogen. Thus, oxygen can be recovered by oxidation of water at the Pt electrode connected to the positive terminal. <u>Write</u> a balanced half-reaction for the oxidation of water.
- **6-3.** When copper is used at both electrodes, gas is generated only at one electrode during the initial stage of electrolysis. <u>Write</u> the half-reaction at the electrode that does not generate gas.

Another species in solution that can be reduced is sodium ion. The reduction of sodium ion to metallic sodium does not occur in aqueous solution, because water is reduced first. However, as Humphrey Davy discovered in 1807, sodium can be made by electrolysis of fused sodium chloride.

**6-4.** Based on these observations, <u>connect</u> the half-reactions with the standard reduction potential (in volts).

Reduction of copper ion (Cu <sup>2+</sup> )	 	+0.340
Reduction of oxygen	•	-2.710
Reduction of water	•	-0.830
Reduction of sodium ion (Na <sup>+</sup> )	•	0.000
Reduction of hydrogen ion	•	+1.230

The electrode potential is affected by other reactions taking place around the electrode. The potential of the Cu<sup>2+</sup>/Cu electrode in a 0.100 M Cu<sup>2+</sup> solution changes as Cu(OH)<sub>2</sub> precipitates. Answer with 3 significant figures for the following problems. The temperature is 25°C. Note that  $K_w = 1.00 \times 10^{-14}$  at 25°C.

- **6-5.** Precipitation of  $Cu(OH)_2$  begins at pH = 4.84. <u>Determine</u> the solubility product of  $Cu(OH)_2$ .
- **6-6.** <u>Calculate</u> the standard reduction potential for  $Cu(OH)_2(s) + 2e^- \rightarrow Cu(s) + 2OH^-$ .
- **6-7.** <u>Calculate</u> the electrode potential at pH = 1.00.

Lithium cobalt oxide and specialty carbon are active ingredients for the positive and negative electrodes, respectively, of a rechargeable lithium battery. During the charge/recharge cycles, the following reversible half-reactions occur.

 $LiCoO_2 \rightleftharpoons Li_{1-x}CoO_2 + x Li^+ + x e^-$ 

 $C + x Li^+ + x e^- \rightleftharpoons CLi_x$ 

The total amount of energy a battery can store is rated in mAh. A battery rated at 1500 mAh can power a device drawing 100 milliamps for 15 hours.

**6-8.** Graphite has lithium intercalation sites between its layers. Assuming a maximum 6:1 carbon-to-lithium intercalation stoichiometry, <u>calculate</u> the theoretical charge capacity of 1.00 gram of graphite to intercalate lithium. Answer in mAh/g with 3 significant figures.

## 7. Hydrogen Economy

Hydrogen is more energy-dense than carbon, by mass. Thus, historically there has been a move toward fuel with higher hydrogen content: coal  $\rightarrow$  oil  $\rightarrow$  natural gas  $\rightarrow$  hydrogen. Cost-effective production and safe storage of hydrogen are two major hurdles to the successful inauguration of a hydrogen economy.

- **7-1.** Consider hydrogen in a cylinder of 80 MPa at 25 °C. Using the ideal gas law, <u>estimate</u> the density of hydrogen in the cylinder in kg/m<sup>3</sup>.
- **7-2.** <u>Calculate</u> the ratio between heat generated when hydrogen is burned and heat generated when the same weight of carbon is burned. The difference comes to a large extent from the fact that the most abundant isotope of hydrogen has no neutron and hydrogen has no inner electron shell.  $\Delta H_{f}^{o}$  [H<sub>2</sub>O(I)] = -286 kJ/mol,  $\Delta H_{f}^{o}$  [CO<sub>2</sub>(g)] = -394 kJ/mol.
- **7-3.** <u>Calculate</u> the theoretical maximum work produced by the combustion of 1 kg hydrogen (a) from the electric motor using hydrogen fuel cell and (b) from the heat engine working between 25 °C and 300 °C. The efficiency (work done/heat absorbed) of an ideal heat engine working between  $T_{cold}$  and  $T_{hot}$  is given by  $[1 T_{cold}/T_{hot}]$ .

$$\begin{split} S^{o}_{298}[H_2(g)] &= 131 \text{ J/(K mol)} \\ S^{o}_{298}[O_2(g)] &= 205 \text{ J/(K mol)} \\ S^{o}_{298}[H_2O(l)] &= 70 \text{ J/(K mol)}. \end{split}$$

If the fuel cell is working at 1 W and the standard potential difference, how long will the electric motor run at what current?

#### 8. Chemistry of Iron Oxides

The nucleus of iron is the most stable among all elements and, therefore, iron accumulates at the core of massive red giant stars where nucleosynthesis of many elements essential for life (such as C, N, O, P, S, etc.) takes place. As a result, among heavy elements iron is quite abundant in the universe. Iron is also abundant on Earth.

**8-1.** Development of a technology for reducing iron oxide to iron was a key step in human civilization. Key reactions taking place in the blast furnace are summarized below.

$C(s) + O_2(g) \rightarrow CO_2(g)$	ΔH° = -393.51 kJ(/mol)	①
$\mathrm{CO}_2(g) + \mathrm{C}(s) \rightarrow 2\mathrm{CO}(g)$	ΔH° = 172.46 kJ(/mol)	2
$Fe_2O_3(s) + CO(g) \rightarrow Fe(s) + CO_2(g)$	ΔH° = ?	③

- 8-1-1. Indicate the reducing agent in each reaction.
- **8-1-2.** <u>Balance</u> reaction ③ and <u>calculate</u> the equilibrium constant of reaction ③ at 1200 °C.  $(\Delta H_{f}^{\circ}(Fe_{2}O_{3}(s)) = -824.2 \text{ kJ/mol}, S^{\circ}(J/mol/K)$ : Fe(s) = 27.28, Fe<sub>2</sub>O<sub>3</sub>(s) = 87.40, C(s) = 5.74, CO(g) = 197.674, CO<sub>2</sub>(g) = 213.74)
- **8-2.** In the manufacture of celadon pottery,  $Fe_2O_3$  is partially reduced in a charcoal kiln to mixed oxides of  $Fe_3O_4$  and FeO. The amount of the different oxides seems to be related to the "mystic" color of celadon ceramics.



 $Fe_3O_4$  (magnetite) itself is a mixed oxide containing  $Fe^{2+}$  and  $Fe^{3+}$  ions and belongs to a group of compounds with a general formula of  $AB_2O_4$ . The oxide ions form a face-centered cubic array. The figure shows the array of oxygens (gray circles) and representative sites for divalent A and trivalent B cations. The dark circle represents a tetrahedral site and the white circle an octahedral site.



**8-2-1.** How many available octahedral sites for iron ions are there in one AB<sub>2</sub>O<sub>4</sub> unit? Certain sites are shared by neighboring units.

 $AB_2O_4$  can adopt a normal- or an inverse-spinel structure. In normal-spinel structure, two B ions occupy two of the octahedral sites and one A occupies one of the tetrahedral sites. In an inverse-spinel structure, one of the two B ions occupies a tetrahedral site. The other B ion and the one A ion occupy octahedral sites.

- **8-2-2.** What percentage of available tetrahedral sites is occupied by either  $Fe^{2+}$  or  $Fe^{3+}$  ion in  $Fe_3O_4$ ?
- **8-2-3** Fe<sub>3</sub>O<sub>4</sub> has an inverse-spinel structure. Draw the crystal field splitting pattern of Fe<sup>2+</sup> and fill out the electrons. The electron pairing energy is greater than the octahedral field splitting.

#### 9. Photolithographic process

Photolithography is a process used in semiconductor device fabrication to transfer a pattern from a photomask to the surface of a substrate. In a typical photolithography process, light is projected, through a mask that defines a particular circuitry, onto a silicon wafer coated with a thin layer of photoresist.

9-1. The earliest photoresists were based on the photochemistry that generates a reactive intermediates from bis(aryl azide). Patterning becomes possible through the cross-linking reaction of the nitrenes generated from the azides.



Bis(aryl azide)

- 9-1-1. <u>Draw</u> two possible Lewis structures of  $CH_3$ -N<sub>3</sub>, the simplest compound having the same active functional group of bis(aryl azide). Assign formal charges.
- 9-1-2. Draw the Lewis structure of nitrene expected from CH<sub>3</sub>-N<sub>3</sub>.
- 9-1-3. <u>Draw</u> the structures for two possible products, when this nitrene from  $CH_3$ -N<sub>3</sub> reacts with ethylene gas ( $CH_2CH_2$ ).
- 9-2. Photoresists consisting of Novolak polymers, utilizes acid to change their solubility. The acid component can be produced photochemically from diazonaphthaquinone. In fact, "Novolaks" have been the representative "positive" photoresists of the modern microelectronic revolution.



When irradiated, diazonaphthaquinone undergoes photochemical decomposition followed by rearrangement eventually producing a carboxylic acid.



9-2-1. <u>Draw</u> three Lewis structures of diazoacetaldehyde (see below), the simplest compound having the same active functional group of diazonaphthaquinone. Indicate formal charges.



#### diazoacetaldehyde

9-2-2. <u>Draw</u> a Lewis structure of the rearranged intermediate, A (see below), generated from diazoacetaldehyde after losing N<sub>2</sub>. A satisfies Lewis' octet rule and reacts with water to form acetic acid, CH<sub>3</sub>CO<sub>2</sub>H.



9-3. Advanced photoresists were invented in 1982 based on chemical amplification. The most popular chemical amplification for positive-tone involves the acid catalyzed deprotection of poly(*p*-hydroxystyrene) resin protected by various acid-sensitive protecting groups such as *t*-butyloxycarbonyl (*t*-BOC).



The thermal decomposition of carbonate ester itself normally occurs well above 150°C.

9-3-1. Two plausible mechanisms have been suggested for this decomposition reaction having relatively high activation energy. <u>Draw</u> expected intermediates and products from this reaction.



9-3-2. In the presence of a trace amount of acid, the reaction temperature can be reduced to below 100°C. <u>Draw</u> expected intermediate F from the following chemical amplification process based on using *t*-BOC.



# 10. Natural Products – Structural Analysis

(9 pts)





Licorice (Glycyrrhizia. Uralensis)

Licorice Root

The flavor extracted from the licorice root is 50 – 150 times sweeter than table sugar.

The most important and abundant compound responsible for the sweetness and medicinal effects of licorice is *glycyrrhizin* ( $C_{42}H_{62}O_{16}$ ).

Glycyrrhizin requires three equivalents of NaOH to effect neutralization.

``When *glycyrrhizin* was subjected to acid hydrolysis, *Glycyrrhizinic acid* (**A**  $(C_{30}H_{46}O_4)$ ) and **B**  $(C_6H_{10}O_7)$  were obtained in a 1:2 molar ratio (figure 1).





When *glycyrrhizin* was methylated with methyl iodide (MeI) at every possible site before hydrolysis, hydrolysis produced **A**' (methyl glycyrrhizinate), **C** and **D** (figure 2). B, C and D exist as mixtures of anomers.

# Figure 2.



Methylation of **C** and **D** with MeI produced the same isomeric mixture of compounds, **J** (figure 3.)

Figure 3.

~	MeI, Ag <sub>2</sub> O	_	MeI, Ag <sub>2</sub> O	
$C (C_9 H_{16} O_7)$		$\mathbf{J} (C_{11}H_{20}O_7)$	◄	$\mathbf{D} (C_{10}H_{18}O_7)$

**C** was reduced with LiAlH<sub>4</sub> to give **K**, and **L** was produced by the reduction of **K**. Oxidative cleavage of vicinal diol of **L** with NalO<sub>4</sub> produced **M** and two equivalents of formaldehyde. Reduction of **M** produced **N**. The structure and stereochemistry of **N** was confirmed by the synthesis of **N** from D-(-)-tartaric acid through methylation followed by reduction (figure 4). A <sup>1</sup>H-NMR spectrum of **L** showed two distinct peaks for methyl groups. (There is no symmetry in L)

Figure 4.



10-1. <u>Complete</u> structures for L , M, and N in the answer sheet.

**10-2.** How many structures for **C** are possible? <u>Complete</u> possible structures for **C**.

To determine the correct structure of **C**, following set of reactions were performed.

**J** was reduced to **E**, and acid hydrolysis of **E** produced **F**. Reduction of **F** generated **G**, and **G** was oxidized with NalO<sub>4</sub> to **H** with formation of one equivalent of formaldehyde. I was obtained from **H** through reduction. Among all

compounds from A to I, only I was optically inactive (figure 5).

# Figure 5



- 10-3. Complete structures for G and I.
- **10-4**. Which one is the correct structure for **C** among ones you have drawn in **10-2**?
- 10-5. <u>Complete</u> structures for **B**, **D**, and **J**.
- **10-6.** <u>Complete</u> the structure for Glycyrrhizin.

#### 11. Enzyme Reaction

Shikimic acid biosynthesis is an important pathway for amino acids, alkaloids and heterocyclic natural product production. Nature converts shikimic acid to chorismic acid through a cascade of enzymatic reactions. Then chorismate mutase catalyzes the conversion of chorismic acid to prephenic acid at the branch point for the biosynthesis of aromatic amino acids such as tyrosine and phenylalanine.



- **11-1.** During the transformation of shikimic acid to chorismic acid, dehydration is occurring. <u>Choose</u> the hydroxyl group in shikimic acid that is lost through above dehydration among all possible reactions.
- **11-2.** Chorismate mutase rearranges chorismic acid into prephenic acid without changing the molecular formula. Chorismic acid becomes prephenic acid through the Claisen rearrangement, a concerted pericyclic process like the Cope rearrangement as shown below:



Based on the following spectral data, <u>propose</u> the structure of prephenic acid.

<sup>1</sup>H-NMR (D<sub>2</sub>O, 250 MHz):  $\delta$  6.01 (2H, d, J = 10.4 Hz), 5.92 (2H, dd J = 10.4, 3.1 Hz), 4.50 (1H, t, J = 3.1 Hz), 3.12 (2H, s). Note that there are three protons, which have been exchanged by D<sub>2</sub>O very fast, and two protons at  $\delta$  3.12, which are exchanged slowly in prephenic acid. <sup>13</sup>C-NMR (D<sub>2</sub>O, 75

MHz):  $\delta$  203, 178, 173, 132 (for two identical carbons), 127 (for two identical carbons), 65, 49, 48.

 $\delta$ , chemical shift; H, integrals; d, doublet; dd, doublet of doublet; J, coupling constant; t, triplet; s, singlet



Chorismate mutase is believed to stabilize the transition state of Claisen rearrangement. Thus it is an interesting target for inhibitor design. Inhibitors, called transition state analog (TSA)s that resemble the transition state (TS, *e.g.*, the species in brackets "[]" above) of the reaction are designed to occupy the active site. Several inhibitors were designed and synthesized, and among them eight turned out to be potent inhibitors of the enzyme. The lower the IC<sub>50</sub> (inhibitor concentration of 50% of the enzymatic activity) value, the better the inhibitor.



- **11-3.** <u>Choose</u> all correct statements based on the structures and  $IC_{50}$  values of above inhibitors. Increase of factor 5 is considered to be important.
  - (a) Configuration of the hydroxyl group plays an important role in the TS and inhibitor design.
  - (b) The presence of both carboxylic groups is important in the TS and inhibitor design.
  - (c) Transition state of the reaction contains two six-membered rings with one chair and one twist-boat conformation.
  - (d) **7** and **8** can be distinguished on the basis of the <sup>1</sup>H-NMR of  $H_a$ .
- **11-4.** <u>Draw</u> the transition state of the transformation of chorismic acid to prephenic acid based on the TSA structures and their  $IC_{50}$  values.
- **11-5.** Compared with the uncatalyzed thermal conversion, chorismate mutase accelerates conversion of chorismic acid to prephenic acid 1.0  $\times$  10<sup>6</sup> fold at 25°C by lowering the activation energy of the reaction. <u>Calculate</u> the decrease in activation energy of chorismate mutase at 25°C.

 $\Delta H^{\sharp}_{uncat}$  is 86,900 J/mol for the thermal conversion of chorismic acid to prephenic acid. At <u>what temperature</u> will the rate of the *uncatalyzed* thermal conversion be the same as that of the *enzyme-catalyzed* conversion at 25°C, assuming that E<sub>a</sub> =  $\Delta H^{\sharp}$ .



			5 g	ots
<b>´   `</b> —	Name:		ma	rks
<b>–</b>	- · · · · · · · · · · · · · · · · · · ·		25	75
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$\mathbf{I}_1$				

1-2.

<b>F</b>				Ę	5 pt	S			
	Name:				marks				
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2-1.			
2-2.			
2-3.			
2-4.			
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<b>2-5.</b>	+ <sup>4</sup> He $\rightarrow$ <sup>(</sup> ) + <sup>1</sup> H					

						5 ;	pts	
́_`_`—	Na	mo.				ma	rks	
_		1	1.		30	30	25	15
$3_{1}$	Sti	udent coc	ie∙					
3-1.								
		residence		time				
		S						
3-2.								
		[						
		(a)	(b)	(c)	(d)		(e	9)
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$\mathbf{J}_1$								
3-4. circle one								
(a)	(b)	(c)	(a, b)	(a, c)			<b>(</b> b,	, c)
(a, b, c)								

T				5	pts		
	Name:			marks			
<b>4</b> 1	Student code:			20 20	20 2	20	
4-1.	Number of atoms Theoretical wt %	H ( (	F ) )	) (	1	)	
4-2.							
4-3.							



				5 pts	
<b>´ `</b> —	Name:		marks		
<b>1</b> 51	Student	code:	20	20	20

5-1.

Species	Concentration
HSO <sub>4</sub>	
SO4 <sup>2-</sup>	
$\mathrm{H}^+$	
OH	
	• •
T-	Name:
---------	---------------
$5_{2}$	Student code:

5-2.

Total volume of 0.80 M NaOH (mL)

T-	Name:
<b>5</b> 3	Student code:

5-3.

T-	Name:
53	Student code:
	Ratio of total aspirin in blood to that in stomach

$T_{-}$	Name:			5 pts	S S
<b>6</b> 1	Student	code:	5 5	5 20 15	5 15 10 10
• -					
6-1.					
6-2.					
6-3.					
6-4.					
	Redu •	ction of oxygen -2.710		•	
	Redu ▪	ction of water -0.830		•	
	Redu •	ction fo sodium ion (Na <sup>+</sup> ) 0.000	•		
	Redu •	+1.230	•		

		5 pts								
· · · —	Name:					ma	rks	6		
-		-								
C	Student	code:	5	5	5	20	15	15	10	10
<b>O</b> 1										
6-5										

-ວ.

T-	Name:	
C	Student	code:
<b>O</b> 2		

6-6.

T- 62	Name: Student	code:
6-7.		

Name: Student	code:
	Name: Student

6-8.







T	Nama		5 pts				
	Name			n	ıark	S	
<b>1 8</b> <sub>1</sub>	Student	– code:	15	25	20	20	20

8-1-1.		
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8-1-2.		
8-2-1.		
8-2-2.		



<b>T</b> –	Name:	5 pts marks
<b>9</b> <sub>1</sub>	Student	code: 15 10 20 15 10 40 10
9-1-1.		
9-1-2		
9-1-3.		



T- 92	Name: Student	code:
9-3-1.		
9-3-2.		









T-	Name:	
<b>10</b> <sub>2</sub>	Student	code:

T-	Name:
<b>10</b> <sub>3</sub>	Student code:



			,	7pts	6	
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Image: Student code: <th>T</th> <th>Namo</th> <th></th> <th>, m</th> <th>7pts nark</th> <th>S KS</th> <th></th>	T	Namo		, m	7pts nark	S KS	
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1   3   1   2     0   0   0   0   0     1   Student code:   1   1   1     11-4   1   1   1   1	-		1	2	1	9	2
1 Student code:	11		0	0	0	0	0
11-4	1	Student code:					
	1-4				<u> </u>		

			7pts				
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11		1 0	3	1 0	2	3 0	
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11-5.		<u>Terrenia</u>					
				1			

T	Name <sup>.</sup>		'n	7pts nark	S IS	
-						
11		1 0	3 0	1 0	2 0	3 0
-	Student code:					

Chemistry for Life,

#### **Chemistry for better Life**



# **Practical Test**



2006. 7. 5 Gyeongsan, Korea

#### **General Directions**

- You have 5 hours to finish the test. Manage your time wisely. You might spend about one hour for Test 1 (10 points), two hours for Test 2 (15 points), and two hours for Test 3 (15 points).
- Write your name and code number on each page of the Answer Sheet.
- There are 7 pages of Test and 7 pages of Answer Sheet.
- Write answers and calculations within the designated box.
- Use only the pen, ruler, and calculator provided
- An English-language version is available.
- Figures to supplement User's Instructions for the spectrophotometer, C-18 cartridge, and pipet filler are provided in a separate sheet.
- Additional samples or supplies will be provided with 1 pt penalty for each item. (except distilled water)
- You may go to the restroom with permission.
- After finishing the test, place all sheets (Test and Answer Sheets) in the envelope and seal.
- Remain seated until instructed to leave the room.
- You may take the pencil case, pen, ruler, calculator, and C-18 cartridges home.

Safety and Disposal

- Wear safety goggles and lab coat.
- No hazardous chemicals are used. All acid, alkali, and dye solutions are dilute. However, it is better to minimize contact with skin. Wipe off with wet Kimwipe in case of contact.
- Do not sniff reagents.
- Dispose used chemicals in the plastic bottle labeled "DISPOSABLE". Discard used test tubes and broken glasses in the "Waste Basket".

### **Apparatus, Chemicals and Supplies**

### Test-1,2 (white basket)

spectropl	notometer		1	
cuvet (1 c	m path-length)		1	
C18 cartr	idge		4	
10 mL sy	ringe		1	
1 mL syri	nge		1	
pasteur p	ipet		3	
1 mL pipe	et		1	
5 mL pipe	et		1	
pipet fille	r		1	
10 mL vo	lumetric flask		2	
buret			1	
test tube			20	
test tube rack			1	
50 mL Erl	enmeyer flask		1	
100 mL b	eaker		2	
silicone b	oulb		2	
three-colo	or pen, ruler		1	
sqeeze bo	ottle		3	
labolad	Solution E	33% etha	nol in water	
labeleu	NaOH solution	less that	n 5 mM	
as	water	distilled	water	
100 mL b	ottle		6	
labeled	Solution R	re in So	d dye Iution E	
as	ed Solution B blue dy in Solutio			
	Solution MD	mixed dy	e of B and R	

Solution MA	mixed acids; acetic acid & salcylic acid in water
КНР	potassium hydrogen phthalate solution
phenolphthalein	0.05% solution

# Test-3 (black basket)

test tube			95	
test tube	rack		1	
spatula			2	
1.5 mL	graduated	pipet	15	
(polyethyle	ene)			
tweezers			1	
pen (for writing on a test tube) 1				
pH test paper 1			1	
100 mL bo	ottle		3	
lah ala d	95% EtOH	95% e	ethanol	
labeled	CH₃CN	acetonitrile		
as	as water distilled water			
30 mL dro	pping bottle		6	
labeled	1M HCI	1M HCI	solution	
as	1M NaOH	1M NaO	H solution	
	•			

	2,4-DNPH	3% 2,4- dinitropl solution	nenylhydrazine
	CAN	20% ceri nitrate s	c ammonium olution
	0.5% KMnO₄	0.5% KM	nO <sub>4</sub> solution
	2.5% FeCl <sub>3</sub>	2.5% Fe	Cl₃ solution
	_		
10 mL via			7
10 mL via		Set⊡ U-′	7
10 mL via		Set⊡ U- <sup>,</sup> Set⊡ U-;	7 1 2
10 mL via		Set⊡ U- Set⊡ U-2 Set⊡ U-3	7 1 2 3
10 mL via		Set⊡ U-' Set⊡ U-' Set⊡ U-' Set⊡ U-'	7 1 2 3 4
10 mL via labeled as		Set □ U Set □ U Set □ U Set □ U Set □ U	7 1 2 3 4 5
10 mL via labeled as		Set U	7 1 2 3 4 5 5 6

The spectrophotometer has three compartments, the light source, the detector, and the cuvet holder. You will find the cover of the cuvet holder open. Leave it open. A cuvet is placed with the label facing the light source (Fig. A). Use this orientation throughout the experiment. The spectrophotometer has been stabilized and is ready for use. Follow the procedure below to take absorbance readings.

- a) Fill the cuvet about 3/4-full with Solution E and insert into the cuvet holder. Do not close the cover of the cuvet holder.
- b) Using the mouse of the computer, move the cursor to REFERENCE and click three times. Then click MEASURE three times and you will get absorbance readings close to zero at ten wavelengths between 470 and 650 nm at 20 nm intervals (Fig. B).
- c) Fill the cuvet with sample solution and click MEASURE three times. You will get absorbance readings for your sample at the same wavelengths. Record absorbance values in the Table in the Answer Sheet.

## How to Use the C18

# Cartridge

- a) The cartridge has an inlet and an outlet (Fig. C). The inlet has a larger diameter.
- b) To wash or elute, first withdraw the liquid with a proper syringe and connect the syringe to the inlet of the cartridge. Then push the liquid slowly with a plunger into the cartridge. (Fig. C & E)
- c) To load the sample, attach the 10 mL syringe to the inlet of the cartridge.

Using a 1 mL pipet, transfer 1.00 mL aliquot of a sample solution to the syringe (Fig. D). Load the sample onto the cartridge using the plunger. Make sure that no amount of sample remains on the syringe. Try to avoid air entering into the cartridge after sample loading.

- d) The cartridge can be reused after washing with Solution E.
- e) Separate the syringe from the cartridge when removing the plunger from the syringe.

### How to Use the Pipet Filler

Move the dial downward to fill the pipet and upward to release the liquid (see Fig. F).

Practical Test 1

#### Reverse-phase Chromatography: Spectrophotometric Analysis

Chromatographic separation followed by spectrophotometric analysis is one of the most widely practiced analytical techniques in chemical laboratories around the world. For example, organic compounds in a complex are often analyzed by reverse-phase mixture liquid chromatography with spectrophotometric detection. In reverse-phase chromatography, hydrophobic interactions between the stationary phase material (usually octadecyl group) and the nonpolar moiety of the analyte is utilized. The chromatogram can be simplified and the compound of interest selectively determined by proper choice of the detector wavelength. In this part of the Practical Test, spectrophotometric analysis of dyes, with and without separation, will be performed.



Food Red No. 40

Methyl Violet 2B

- 1-1. Spectrophotometric Analysis of R and B in a Mixed Solution
  - a) Measure absorbance of both Solutions R (3.02 x 10<sup>-5</sup> M) and B (1.25 x 10<sup>-5</sup> M) (Fig. A & B). <u>Fill in</u> the Table in the Answer Sheet with your measurements. <u>Draw</u> absorption spectra for the red dye in red ink and for the blue dye in blue ink (Fig. 1-1).
  - b) Repeat absorbance measurements for Solution MD. Solution MD is a mixture of Solution R and B in a certain ratio. <u>Add</u> the spectrum in black ink to Fig. 1-1.
  - c) Based on the Beer-Lambert law, <u>determine</u> the molar concentration of both dyes in Solution MD using the data in the Table. Do not determine the fraction of one dye by subtracting the fraction of another dye from 1.

1-2. Chromatographic Separation Followed by Spectrophotometric Analysis

- a) Elute the cartridge with about 10 mL of Solution E using 10 mL syringe (Fig. C).
- b) Load 1.00 mL of solution MD onto the cartridge (Fig. D).

- c) Using 1 mL syringe, elute with Solution E (Fig. E). Collect the solution eluting through the outlet in a 10 mL volumetric flask. Repeat until the red compound is completely eluted and collected.
- d) Fill the flask to the 10 mL mark with Solution E and mix. Call this Solution F.
- e) Obtain the absorption spectrum of solution F as in Experiment 1-1. Dilution takes place during elution. Therefore, multiply the measured absorbance by 10 when drawing the spectrum for Solution F. <u>Draw</u> spectrum with broken line in Fig. 1-1 in red ink.
- f) Dilute Solution R as necessary and construct a calibration curve, at a wavelength of your choice, for analysis of the red dye (R) in Solution F. <u>Draw</u> a calibration curve in the answer sheet (X-axis, concentration; Y-axis, absorbance, Fig. 1-2). Indicate the wavelength used. The calibration curve must have three points in addition to the origin. <u>Mark</u> the position of Solution F on the calibration curve.
- g) <u>Report</u> the concentration of R in the original Solution MD.
- h) Compare this concentration with the value you obtained in Experiment 1-1 and <u>report</u> the recovery (amount eluted/amount loaded) associated with chromatography.

**Practical Test 2** 

#### Reverse-phase Chromatography: Acid-Base Titration of Acetic Acid and Salicylic Acid

Acetic acid (AA) and salicylic acid (SA) are slightly different in polarity and thus can be separated on a reverse-phase cartridge using distilled water as eluent. AA is eluted first. The total amount of AA and SA in a mixed solution will be determined by titration. Then, AA and SA will be separately determined following chromatographic separation.

2-1. Determination of the Total Amount of AA and SA in a Mixed Acid (MA) Solution

- a) Titrate 10 mL of distilled water with the NaOH (< 5 mM) solution provided. <u>Report</u> blank acidity in 1 mL of distilled water in terms of the volume of the NaOH solution. Take this blank acidity into account for all solutions in subsequent data analyses. Show corrections in the calculation part in the answer sheet.
- b) Standardize NaOH solution with 2.00 mL of the standard KHP (potassium hydrogen phthalate) solution (1.00 x 10<sup>-2</sup> M) provided. Repeat and <u>report</u> the concentration of the NaOH solution. <u>Show</u> how you accounted for the blank acidity.
- c) Withdraw 1.00 mL of Solution MA and determine the total acidity. Repeat and <u>report</u> the total number of moles of AA and SA combined in 1.00 mL of Solution MA.
- 2-2. Reverse-phase Separation and Titration
  - a) Elute a new C-18 cartridge with about 10 mL of distilled water using 10 mL syringe.

- b) Load 1.00 mL of Solution MA onto the cartridge. Collect the liquid eluting at the outlet in tube 1 (Fraction 1).
- c) Elute with 1 mL of distilled water. Collect the eluent in a test tube (Fraction 2). Repeat until Fraction 20 is collected. You will have 20 test tubes with about 1 mL liquid in each tube.
- d) Titrate acidity in each test tube. <u>Report</u> volume of the NaOH solution consumed and the amount of acid(s) in each test tube. <u>Make</u> a graph in the answer sheet (Fig. 2-2) showing the amount of acid(s) in each test tube.
- e) Blank acidity and the background (due to leaching out of residual materials from the column) must be subtracted. In determining the amount of eluted AA, disregard tubes containing only trace amounts of acids. Tube 2 and 3 contain most AA. <u>Calculate</u> the total amount of AA eluted by adding the amount of AA in tubes. Similarly <u>calculate</u> the total amount of SA eluted. Indicate, in Fig. 2-2, which fractions you used to get the amount of each acid.
- f) Calculate the mole percent of AA in solution MA.
**Practical Test 3** 

## **Qualitative Analysis of Organic Compounds**

In this experiment your task is to identify seven solid unknowns from the list of compounds on page 7 that are common drugs in everyday life and valuable agents in organic chemistry. To achieve this, perform chemical tests on unknowns according to the following procedures and analyze your results.

- Unknowns Labeled Set
U-1, Set
U-2, Set
U-3, Set
U-4, Set
U-5, Set
U-6, Set
U-7

### **Procedure**

**Helpful Comments** 

- a) The weight of a spatula tip-full of a solid is about 15~20 mg.
- b) Wipe spatula cleanly with Kimwipe between uses.
- c) After adding any reagent described below to a solution of an unknown sample, mix the contents thoroughly and observe the resulting mixture carefully.
- d) To get full marks, you should perform all the tests and fill out the table.

Test 1: Solubility test

To a test tube, add a spatula tip-full (15~20 mg) of an unknown sample and 1 mL of  $CH_3CN$ . Shake the test tube and report the solubility. Repeat the test with 1M HCI, water, and 1M NaOH.

### Test 2: 2,4-DNPH test

Place about 15~20 mg of an unknown sample in a test tube and dissolve with 2 mL of 95% EtOH (For the water soluble unknowns, dissolve about 15~20 mg of an unknown in 1 mL of water). Add five drops of the 2,4-dinitrophenylhydrazine solution in concentrated sulfuric acid and 95% ethanol (labeled as 2,4-DNPH).

### Test 3: CAN test

Mix 3 mL of the cerium(IV) ammonium nitrate solution in dilute  $HNO_3$  (labeled as CAN) with 3 mL of CH<sub>3</sub>CN in a test tube. In another test tube add about 15~20 mg of an unknown sample in 1 mL of the mixed solution. (For the water soluble unknown samples, dissolve about 15~20 mg of an unknown sample in 1 mL of water first, and then add 1 mL of CAN.) If there is a color change in the solution, the solution may contain alcohol, phenol or aldehyde.

### Test 4: Baeyer test

In a test tube, dissolve about  $15\sim20$  mg of an unknown sample in 2 mL of CH<sub>3</sub>CN (For the water soluble unknown samples, dissolve about  $15\sim20$  mg of an unknown in 1 mL of water). To the solution, slowly add five drops of the 0.5% KMnO<sub>4</sub> solution, drop by drop while shaking.

### Test 5: pH test

In a test tube, dissolve about 15~20 mg of an unknown sample in 2 mL of 95% EtOH (For the water soluble unknown samples, dissolve about 15~20 mg of an unknown sample in 1 mL of water). Measure the pH of the solution with pH paper.

### Test 6: Iron(III) chloride test

Take the solution from Test 5 and add five drops of a 2.5% FeCl<sub>3</sub> solution.

### **Results**

1. Record your test results in the answer sheet. Write *O* if soluble and *X* if insoluble for the solubility tests. Write (+) for the positive reactions and (–) for the negative reactions for tests  $2 \sim 4$  and 6. Write *a*, *b* and *n* for acidic, basic or neutral, respectively, for pH test 5.

2. Based on your test results, identify the most plausible structures for the unknown compounds from the provided list of compounds. Write the compound initial in appropriate box.



### Possible Unknown Compounds

<b>P</b> –	Name:					1( m	) p ar	ots ks				
<b>1</b> 1	Student		code:	3	9	2 5	3	2 5	3 0	5		
Circle the Set MD & MA you	Number of Solution received.	1	2			3			4		5	
1-1. Spectrophotometric Analysis of R and B in a												
a), b) Absorbar	nce measurements (3 marks)											

Movelongth (nm)		Absorbance	
wavelength (nin)	Solution R	Solution B	Solution MD
470			
490			
510			
530			
550			
570			
590			
610			
630			
650			

# Fig. 1-1Absorption spectra of dyes (9 marks)



P-	Name:
12	Student code:

c) Concentration of R & B (Beer-Lambert law) (25 marks)

Concentration of R in solution MD	М
Concentration of B in solution MD	М

P-	Name:
13	Student code:

# 1-2. Chromatographic Separation Followed by Spectrophotometric Analysis e) solution F (3 marks)

Wavelength (nm)	470	490	510	530	550	570	590	610	630	650
absorbance										

### f) Calibration curve (25 marks)

Concentration (M)		
Absorbance at ( ) nm		

### Fig. 1-2 Calibration curve

++++	++++	 	++++	++++	++++	++++	++++	++++	++++	++++	+++++	++++	+++++	++++	++++	+++++	++++	+++++	
++++					++++														
++++					++++	++++		++++					++++	++++	++++				

### g) Concentration of R in solution MD (30 marks)

h) Percent recovery of R (5 marks)

4

Μ

					15	pts				
P-	Name:				ma	rks				
<b>1</b> 21	Student	code:	5	10	10	25	30	20		
2-1. Determinati	2-1. Determination of the Total Amount of AA and SA in a Mixed Solution									

# a) blank titration (5 marks)

Volume of NaOH solution consumed for blank titration	mL
NaOH solution equivalent to blank acidity of 1 mL water	mL

## b) standardization of the NaOH solution (10 marks)

KHP used (mL)	NaOH consumed (mL)	concentration of NaOH (M)				
		Concentration measured				

## c) total amount of AA and SA in 1.00 mL of Solution MA (10 marks)

Solution MA used (mL)	AA and SA in 1 mL MA (mol)					
		Total amount measured				

ſ					15	pts		
P-	Name:				ma	rks		
<b>1</b> <b>2</b> 1	Student	code:	5	10	10	25	30	20

P-	Name:
$2_{2}$	Student code:

### 2-2. Reverse-phase Separation and Titration

d) Amount of acid(s) in each test tube (25 marks)

tube	1	2	3	4	5	6	7
NaOH (mL)							
Acid (mmol)							
tube	8	9	10	11	12	13	14
NaOH (mL)							
Acid (mmol)							
tube	15	16	17	18	19	20	
NaOH (mL)							
Acid (mmol)							

Fig. 2-2 Amount of acid(s) in each test tube



P-	Name:
<b>2</b> <sub>3</sub>	Student code:

## e) amount of AA and SA eluted (30 marks)

Fraction numbers used	
Total amount of AA eluted (1 <sup>st</sup> peak)	mmol
Total amount of SA eluted (2 <sup>nd</sup> peak)	mmol

## f) mole percent of AA in Solution MA (20 marks)

AA from e)	mmol
AA + SA from c)	mmol

		Name	): 				15	pts
P.	-						marks	
3	1	Student code:				31.5	70	
Circle vo	ur Set.							
		A F	B G	С Н	D		E	
3-1 Test I	Result (31	.5 marks	)					
		_						
unl test	knewn	U-1	U-2	U-3	U-4	U-5	U-6	U-7
	CH₃CN							
0-1-1 ***	1M HCI							
Solubility	water							_
	1M NaOH							
Test 2	2 (DNPH)							
Test	3 (CAN)							
Test 4	(KMnO₄)							
Tes	t 5 (pH)							
Test	6 (FeCl₃)							
3-2 Identi	ty of unkr	iown cor	npound	ds (70 mark	(s)	1		
U-1	U-2	U	-3	U-4	U-5		U-6	U-7

D	Name:		15	pts
	T tunne.	marks		
<b>-</b> <b>3</b> 1	Student	code:	31.5	70